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1960s, the essential  
features of the subject  
were well understood,  
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kinetics of reactive radicals were rapidly accumulating. This short text sets out to present that basic understanding in a modern context, in which extensive use of radical reactions is now being made in organic synthesis, and where, in the life sciences, reactive radicals are being recognised both as mediators of ...

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By the 1960s, the

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essential features of radical chemistry were understood. This work presents that understanding, in which the use of radical reactions is being made in organic synthesis, and in the life sciences, where reactive radicals are recognised as mediators of many disease conditions, and as key players in mechanisms of enzyme action.



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polymerization,  
polymer reaction  
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polymer  
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Author Bios. Michael  
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of sulfinates and 2H-  
indazoles to sulfonyl  
radical and radical  
cationic 2H-indazoles,  
facilitating the direct  
radical-radical ...

## **Electrochemical Radical-Radical Cross-Coupling Approach ...**

Radical, also called  
Free Radical, in  
chemistry, molecule  
that contains at least  
one unpaired electron.  
Most molecules contain

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even numbers of electrons, and the covalent chemical bonds holding the atoms together within a molecule normally consist of pairs of electrons jointly shared by the atoms linked by the bond.

## **Radical | chemistry | Britannica**

In chemistry, a radical (more precisely, a free radical) is an atom, molecule, or ion that

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has unpaired valence electrons or an open electron shell, and therefore may be seen as having one or more "dangling" covalent bonds.

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2.1. Initiation. This first step is represented by the general following reaction, where M is the reactant molecule, R the radical species

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and  $K_1$  the reaction  
rate:  $(1) M \rightarrow k_1 R \cdot$ .

Firstly, the lowest-energy configuration of the dioxygen molecule ( $O_2$ ), which is a stable, relatively unreactive diradical in a triplet spin state, is forced into a spin-paired state, or singlet oxygen ( $^1O_2$ ).

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A radical is a chemical

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entity with an unpaired electron in it. A radical can be electrically neutral, radical cation charged positively, or radical anion charged negatively. The

Radicals Examples. 1. In ultraviolet light the chlorine molecule  $\text{Cl}_2$  undergoes homolysis to form two radicals of  $\text{Cl}$ .

**What are Free Radicals? - Definition,**



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**Examples, Types ...**

Hydroxyl Radical's Role  
in the Remediation of a  
Common Herbicide, 2,4  
-Dichlorophenoxyacetic  
Acid (2,4-D). The  
Journal of Physical  
Chemistry A 2004 , 108  
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Chemistry Primers) by

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helps upper

undergraduates

understand the basics

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A 20% yield (!) of new knot-shaped molecules with amide structure is

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surprisingly obtained from simple building blocks in a one-step synthesis. The molecular structure is confirmed by X-ray structure analysis. The centrosymmetric unit cell and attempts at chromatographic enantiomeric separation on chiral column material reveals that the trefoil knot (see picture) is, as expected ...

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## **A Molecular Knot with Twelve Amide Groups—One-Step ...**

For radicals within larger molecules, see Moiety (chemistry). The hydroxyl radical, Lewis structure shown, contains one unpaired electron. In chemistry, a radical is an atom, molecule, or ion that has an unpaired valence electron. With some exceptions, these unpaired electrons make radicals highly

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chemically reactive.

## **Radical (chemistry) - Wikipedia**

Radical (chemistry), an atom, molecule, or ion with unpaired valence electron(s) Radical surgery, where diseased tissue or lymph nodes are removed from a diseased organ; Mathematics. Radical expression involving roots, also known as an  $n$ th root; Radical

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symbol ( $\sqrt{\quad}$ ), used to indicate the square root and other roots

## **Radical - Wikipedia**

Beyond the realms of mere curiosity and fundamental research, the commercialization of organic electronics [e.g., organic light-emitting diodes (OLEDs)] has boosted the impetus to elucidate the fundamentals regarding the vast

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chemistry and related physical properties and device opportunities of organic materials (Fig. 1.1) [1].

**An Introduction to  
Radical Polymers |  
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radical reactions and explains the principles of ion radical organic chemistry. Methods of determining ion-radical mechanisms and controlling ion radical reactions are also



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reviewed. When applicable, issues relating to ecology and biology are addressed. The inorganic participants in the ion radical organic reactions are also considered.

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